

Please write clearly in block capitals.

Centre number

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Candidate number

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Surname

Forename(s)

Candidate signature

GCSE COMBINED SCIENCE: TRILOGY

H

Higher Tier
Chemistry Paper 1H

Thursday 16 May 2019

Morning

Time allowed: 1 hour 15 minutes

Materials

For this paper you must have:

- a ruler
- a scientific calculator
- the periodic table (enclosed).

Instructions

- Use black ink or black ball-point pen.
- Fill in the boxes at the top of this page.
- Answer **all** questions in the spaces provided.
- Do all rough work in this book. Cross through any work you do not want to be marked.
- In all calculations, show clearly how you work out your answer.

Information

- The maximum mark for this paper is 70.
- The marks for questions are shown in brackets.
- You are expected to use a calculator where appropriate.
- You are reminded of the need for good English and clear presentation in your answers.

For Examiner's Use	
Question	Mark
1	
2	
3	
4	
5	
6	
TOTAL	



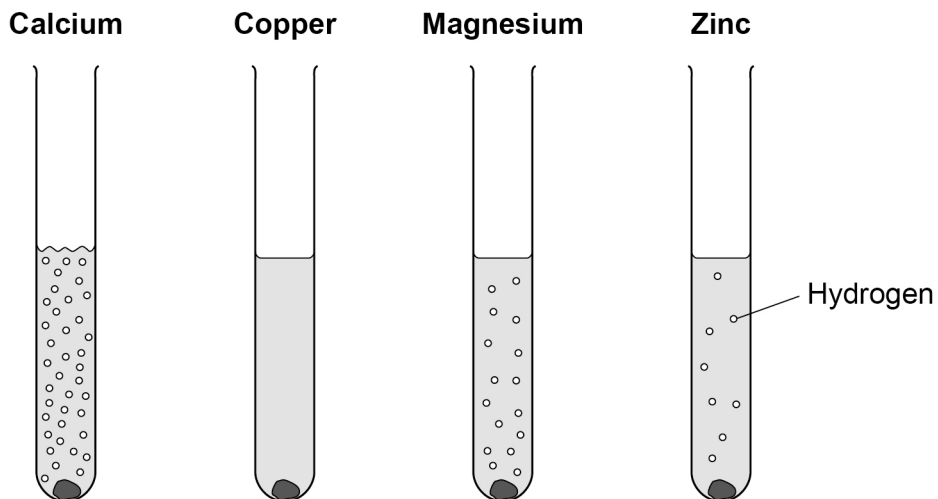
J U N 1 9 8 4 6 4 C 1 H O 1

0 1

This question is about reactions of metals.

Figure 1 shows what happens when calcium, copper, magnesium and zinc are added to hydrochloric acid.

Figure 1



0 1 . 1

What is the order of decreasing reactivity of these four metals?

[1 mark]

Tick (✓) **one** box.

Zn Ca Cu Mg

☐

Ca Cu Mg Zn

☐

Cu Zn Ca Mg

☐

Ca Mg Zn Cu

☐


A student wants to make a fair comparison of the reactivity of the metals with hydrochloric acid.

0 1 . 2

Name **two** variables that must be kept constant.

[2 marks]

1 _____

2 _____

0 1 . 3

What is the independent variable in this reaction?

[1 mark]

0 1 . 4

Predict the reactivity of beryllium compared with magnesium.

Give a reason for your answer.

Use the periodic table.

[2 marks]

Reason _____

0 1 . 5

A solution of hydrochloric acid contains 3.2 g of hydrogen chloride in 50 cm³

Calculate the concentration of hydrogen chloride in g per dm³

[3 marks]

Concentration = _____ g per dm³



0 2

This question is about salts.

Ammonium nitrate solution is produced when ammonia gas reacts with nitric acid.

0 2 . 1

Give the state symbol for ammonium nitrate solution.

[1 mark]

0 2 . 2

What is the formula of nitric acid?

[1 mark]Tick (✓) **one** box.

HCl

☐HNO₃☐H₂SO₄☐NH₄OH☐**0 2 . 3**

Ammonia gas dissolves in water to produce ammonia solution.

Ammonia solution contains hydroxide ions, OH⁻

A student adds universal indicator to solutions of nitric acid and ammonia.

What colour is observed in each solution?

[2 marks]

Colour in nitric acid

Colour in ammonia solution



0 2 . 4

The student gradually added nitric acid to ammonia solution.

Which row, **A**, **B**, **C** or **D**, shows the change in pH as the nitric acid is added until in excess?

[1 mark]

Tick (✓) **one** box.

	pH of ammonia solution at start	pH after addition of excess nitric acid	
A	10	7	☐
B	2	10	☐
C	7	1	☐
D	10	2	☐

0 2 . 5

Calculate the percentage by mass of oxygen in ammonium nitrate (NH_4NO_3).

Relative atomic masses (A_r): H = 1 N = 14 O = 16

Relative formula mass (M_r): $\text{NH}_4\text{NO}_3 = 80$

[3 marks]

Percentage by mass of oxygen = _____ %

Question 2 continues on the next page

Turn over ►



Describe a method to investigate how the temperature changes when different masses of ammonium nitrate are dissolved in water.

You do **not** need to write about safety precautions.

[6 marks]

[illegible]

14



Turn over for the next question

*Do not write
outside the
box*

**DO NOT WRITE ON THIS PAGE
ANSWER IN THE SPACES PROVIDED**

Turn over ►



0	3
---	---

This question is about oxygen.

0	3	.	1
---	---	---	---

Hydrogen reacts with oxygen.

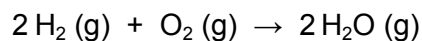
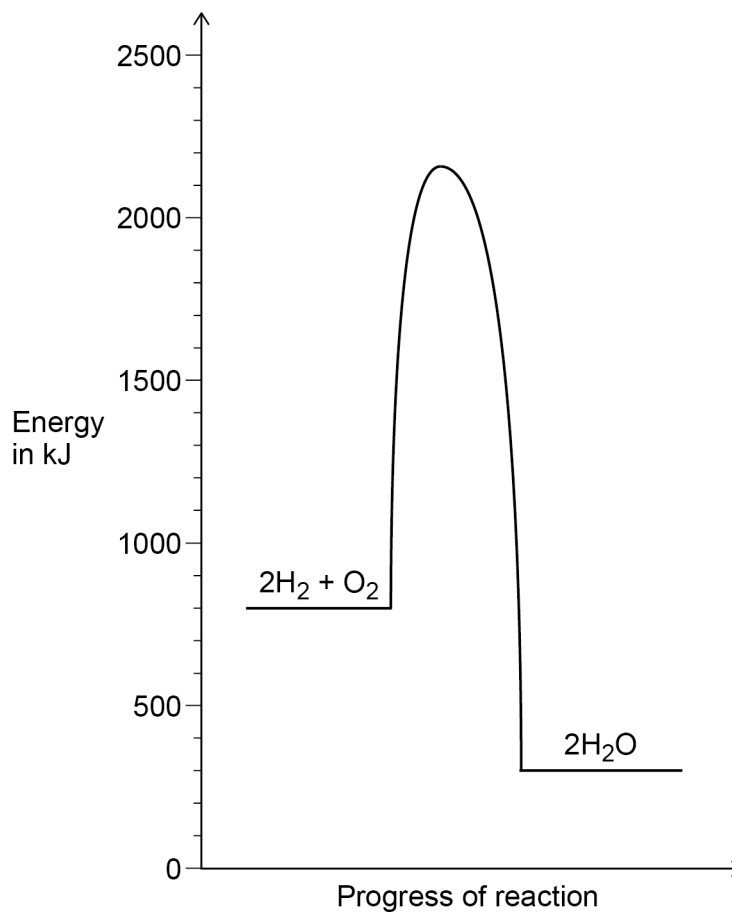


Figure 2 shows the relative energies of the reactants and products at a certain temperature.

Figure 2

Label the activation energy on **Figure 2**.**[1 mark]**

0 3 . 2

Determine the overall energy change for the reaction between hydrogen and oxygen shown in Question **03.1**

Use **Figure 2**.

[2 marks]

Energy change = _____ kJ

0 3 . 3

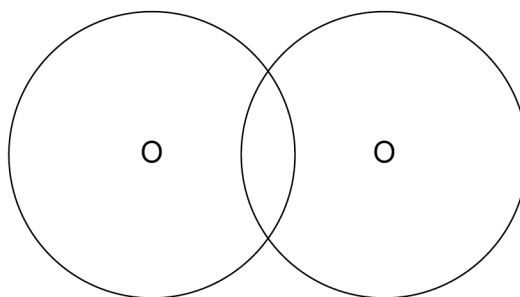
Oxygen is in Group 6 of the periodic table.

Figure 3 shows the outer energy levels in one molecule of oxygen (O_2).

Draw the electrons in the outer energy levels in **Figure 3**.

[2 marks]

Figure 3



Question 3 continues on the next page

Turn over ►



0 3 . 4

The equation shows the decomposition of hydrogen peroxide.

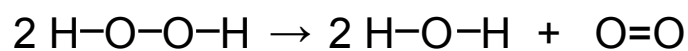


Table 1 shows the bond energies.

Table 1

Bond	O–O	O=O	O–H
Bond dissociation energy in kJ per mole	138	496	463

Calculate the overall energy change for the reaction.

[3 marks]

Energy change = _____ kJ



0	4
---	---

This question is about elements in the periodic table.

0	4	.	1
---	---	---	---

What order did scientists use to arrange elements in early periodic tables?

[1 mark]

0	4	.	2
---	---	---	---

In the early periodic tables some elements were placed in the wrong groups.

Mendeleev overcame this in his periodic table.

Give **one** way Mendeleev did this.

[1 mark]

Question 4 continues on the next page

Turn over ►



Table 2 shows the boiling points of fluorine, chlorine and bromine.

Table 2

Element	Boiling point in °C
Fluorine	−186
Chlorine	−34
Bromine	+59

0 4 . 3

Explain why the boiling points in **Table 2** are low.

[2 marks]

0 4 . 4

Explain the trend in the boiling points in **Table 2**.

[3 marks]



0 4 . 5 Explain why neon is unreactive.

Give the electronic structure of neon in your answer.

[2 marks]

0 4 . 6 How many atoms are there in 1 g of argon?

The Avogadro constant is 6.02×10^{23} per mole.

Relative atomic mass (A_r): Ar = 40

[2 marks]

Number of atoms in 1 g = _____

11

Turn over for the next question

Turn over ►



0	5
---	---

This question is about electrolysis.

0	5	.	1
---	---	---	---

Some metals are extracted from molten compounds using electrolysis.

Why is electrolysis used to extract some metals?

[1 mark]

0	5	.	2
---	---	---	---

Aluminium is produced by electrolysis of a molten mixture.

What **two** substances does the molten mixture contain?

[2 marks]

1

2

0	5	.	3
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Copper and chlorine are produced when molten copper chloride is electrolysed.

Complete the half equation for the reaction at each electrode.

[2 marks]

Half equation at negative electrode



Half equation at positive electrode

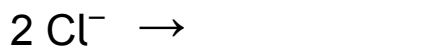
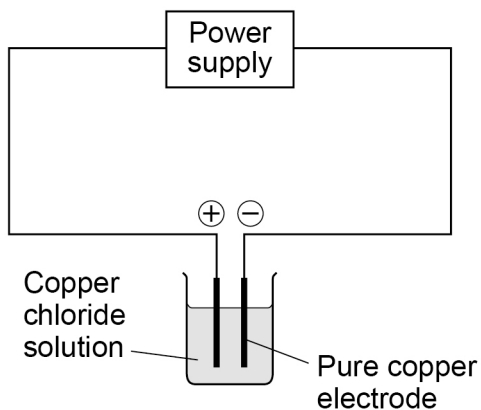


Figure 4 shows the apparatus a student used to electrolyse copper chloride solution.

Figure 4



The student:

- measured the mass of copper deposited on the negative electrode after 60 minutes
- compared the mass deposited with the expected value.

0 5 . 4

Suggest **two** reasons why the mass deposited was different from the expected value.

[2 marks]

1 _____

2 _____

Question 5 continues on the next page

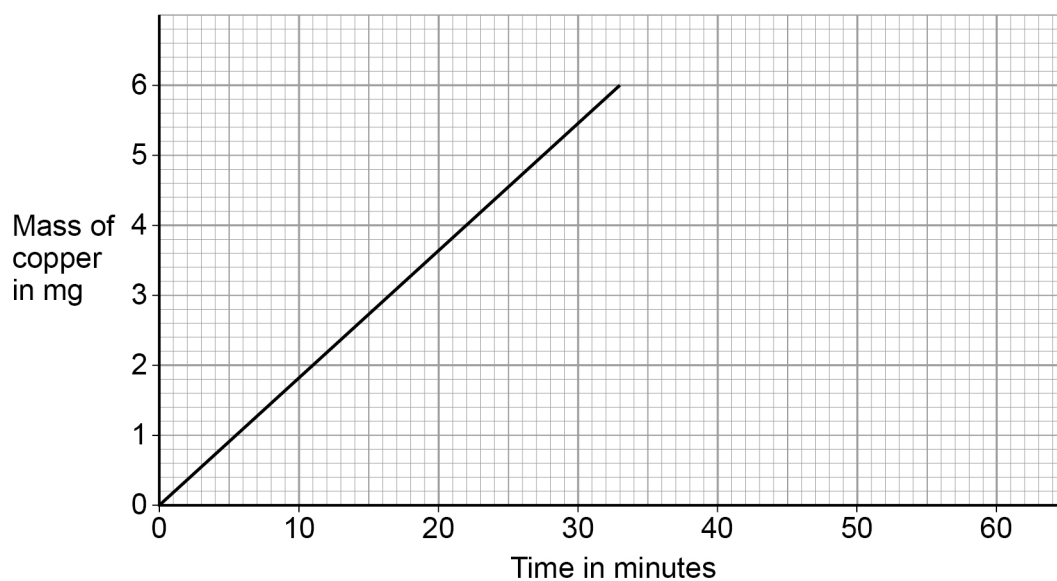
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0 5 . 5

Figure 5 shows the expected mass of copper produced each minute.

Figure 5



Determine the expected mass of copper after 24 hours.

Use **Figure 5**.

[3 marks]

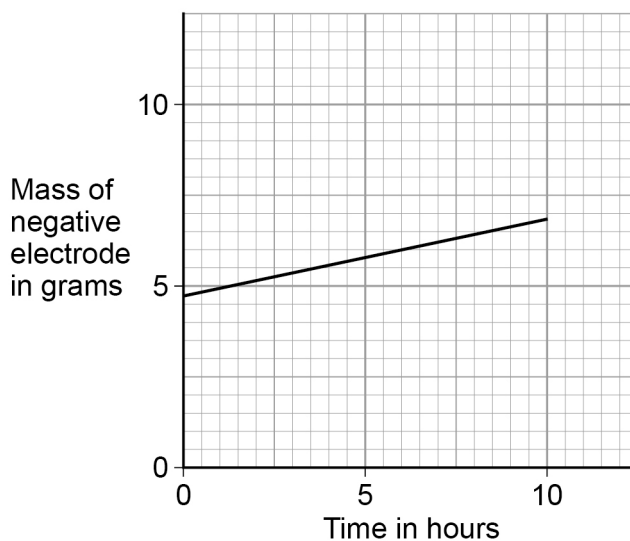
Mass = _____ mg



Silver nitrate solution is electrolysed.

Figure 6 shows the change in mass of the negative electrode over 10 hours.

Figure 6



0 5 . 6 Determine the mass of the negative electrode at the start of the experiment.

Use **Figure 6**.

[1 mark]

0 5 . 7 Calculate the gradient of the line in **Figure 6**.

Give the unit.

[3 marks]

Gradient _____

Unit _____



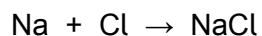
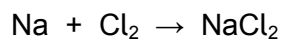
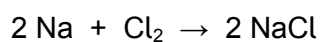
0 6

This question is about sodium.

0 6 . 1

Sodium reacts with chlorine.

What is the balanced equation for the reaction?

[1 mark]Tick (✓) **one** box.☐☐☐☐**0 6 . 2**

Hot sodium is put in a gas jar of chlorine.

Describe the observations made before, during and after the reaction.

[3 marks]

Before reaction _____

During reaction _____

After reaction _____



0	6	.	3
---	---	---	---

Explain why sodium is less reactive than potassium.

[4 marks]

Question 6 continues on the next page**Turn over ►**

[6 marks]

[illegible]

14

